**Practice Exam 5**

**Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period\_\_\_\_\_\_\_ Score\_\_\_\_\_\_**

**Name the following:**

**1a. N3O4 Answer: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**1b. P2Cl5 Answer: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**2a. BrF6 Answer: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**2b. S2O8 Answer: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Write the formulas for the following:**

**3a. Diiodine pentaoxide Answer: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**3b. carbon trichlorideAnswer: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**4. Draw the Lewis structure C5H12**

**Include the electron inventory.**

**5. Draw the Lewis structure for HBr .**

**Include the electron inventory.**

**6. Draw the Lewis H2S**

**Include the electron inventory.**

**7. Draw the Lewis structure forNH4+**

**Include the electron inventory.**

**8. Draw the Lewis structure forCO2**

**Include the electron inventory.**

**9. Describe the molecular geometry (shape) of the molecule in**

**a. problem 5 \_\_\_\_\_\_\_\_\_\_\_**

**b. problem 6 \_\_\_\_\_\_\_\_\_\_\_**

**c. problem 7 \_\_\_\_\_\_\_\_\_\_\_**

**d. problem 8\_\_\_\_\_\_\_\_\_\_\_**

**10. Is the molecule polar or nonpolar in**

**a. problem 5 \_\_\_\_\_\_\_\_\_\_\_**

**b. problem 6 \_\_\_\_\_\_\_\_\_\_\_**

**c. problem 7 \_\_\_\_\_\_\_\_\_\_\_**

**d. problem 8\_\_\_\_\_\_\_\_\_\_\_**

**11. Draw the Lewis structure forSO42-**

**Include the electron inventory.**